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Chapter 12 Stoichiometry . Stoichiometry . A balanced chemical equation provides the same kind of quantitative information that a recipe does. Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. Stoichiometry .

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Chapter 11 is a review of reaction types and balancing chemical reactions. There is also an introduction to predicting the products of a chemical reaction. Chapter 12 is an introduction to

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stoichiometry. This includes doing mass-to-mass and volume-to-volume calculations, percent yield and identification of limiting reagents.

Unit 5 - Chapters 11 & 12 - Mrs. Gingras' Chemistry Page

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Zumdahl's Chapter 12 Chemical Kinetics ... A B Reaction Rate Stoichiometry requires $d[A]/dt = -d[B]/dt$ But $d[A]/dt$ can itself be a function of time. It falls rapidly initially. Then it approaches its equilibrium value, as $[A]$ on the graph, asymptotically.

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